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(54) Title: METHOD FOR COATING GLASS SUBSTRATES		
(57) Abstract A method for coating glass by chemical-vapor deposition below 200 °C, at atmospheric pressure of a gaseous mixture of a tin oxide precursor, a silicon dioxide precursor and an accelerant such as triethyl phosphite provides deposition rates greater than 350 Å per second. The layer of material deposited can be combined with other layers to produce an article with specific properties such as controlled emissivity, refractive index, abrasion resistance, and appearance.		

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METHOD FOR COATING GLASS SUBSTRATES

BACKGROUND OF THE INVENTION

5 *Cross-Reference to Related Application.* This application is a continuation-in-part of our copending United States Patent Applications, Serial Numbers 07/814,366, filed 12/26/91 and 07/814,352, filed 12/27/91.

Field of the Invention. The present invention is in the field of methods for coating on substrates. More particularly, this invention is in the field of
10 methods for depositing coatings at enhanced deposition rates on glass or glass articles to provide controlled refractive index, improved emissivity characteristics and/or appearance and abrasion resistance, and to complement or enhance other properties.

Description of the Prior Art. Transparent semi-conductor films such as indium
15 oxide, cadmium stannate, or doped tin oxide, can be applied to various transparent substrates such as, e.g., soda-lime glasses, in order to reflect long-wavelength infrared radiation. Transparent dielectric films such as titanium dioxide or undoped tin oxide can be applied to transparent articles such as glass bottles to form a base coat for a second coating with a specific function.
20 Depending on the thickness of the semiconductor or dielectric film, various reflected iridescent colors may be observed. This iridescent effect is considered to be detrimental to the appearance of the glass in applications such as windows with low emissivity, or bottles for food or beverages.

Methods and apparatus for coating glass, and especially continuous
25 coating on moving glass, are known in the art. A description of apparatus useful in preparing a coated-glass article is found in Lindner, U.S. 4,928,627, made a part of this disclosure by reference herein.

Various procedures have been devised for reducing or eliminating iridescence. For the low-emissivity application, Zaromb, in US 3,378,396,
30 describes an article comprising a transparent glass substrate coated with tin and silicon oxides; the coating varies gradually in composition from a high ratio of silicon oxide to tin oxide at the substrate surface, gradually changing to almost pure tin oxide, and changing further to a ratio of not more than

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60% silicon oxide to not less than 40% tin oxide at the interface of that coating with the atmosphere. The refractive index of the coating nearest to the substrate is about 1.5, substantially the refractive index of silica glass, and changes to about 2.0, the refractive index of tin oxide, at the air interface, providing an intermediate coating layer without an optical interface. The article so coated has little to no iridescence in reflected light. Zaromb teaches that aqueous solutions of tin and silicon chlorides can be spray-applied to achieve his coatings. Spray applications are usually batch operations which do not yield high-quality, uniform films; there is no mention of other means of application such as chemical-vapor deposition (CVD). He also fails to give any indication of the deposition rate, a key parameter for a commercial industrial application.

Another approach is described by Gordon in US 4,187,336. One or more layers of a transparent material with a refractive index intermediate between that of a glass substrate and a conductive tin oxide film are deposited by atmospheric-pressure CVD between the glass and the tin oxide film. It is necessary for the intermediate layers to have specific refractive indices and thicknesses in order to be effective. It is noted that when the intermediate films contained silicon dioxide, suitable volatile compounds were found to be silane, dimethylsilane, diethylsilane, tetramethyl silane, and the silicon halides. No other precursors are mentioned. The deposition rates obtained for the processes described were on the order of from 10 to 20 Ångstroms per second (Å/sec.). Such rates are an order of magnitude below those necessary for a commercial industrial process.

In United States Patent 4,206,252, Gordon describes a process for depositing mixed oxide and nitride coating layers of continuously varying refractive index between a glass substrate and an infra-red-reflecting coating, whereby the film iridescence is eliminated. When silicon dioxide is part of the mixed oxide film, the patent teaches that volatile silicon compounds with Si-Si and Si-H bonds are suitable precursors. Compounds such as 1,1,2,2-tetramethyldisilane, 1,1,2-trimethyldisilane, and 1,2-dimethyldisilane are

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disclosed. All of the compounds containing Si-Si and Si-H bonds to which reference is made are expensive, and none are commercially available.

In U.S. 4,386,117, Gordon describes a process for preparing mixed silicon oxide/tin oxide coatings at specific refractive indices or a continuous gradient as taught by Zaromb in US 3,378,396, at optimum deposition rates of 80 to 125 Å/sec, using alkoxy-peralkylpolysilane precursors such as methoxypentamethyldisilane or dimethoxytetramethyldisilane. Again, the silica precursors cited and inferred are impractical for industrial use, because none of them is commercially available on a large scale.

Lagendijk, in United States Patent 5,028,566, notes in column 4 that tetraethylorthosilicate (TEOS) suffers from a number of disadvantages in its application to a substrate by low-pressure CVD; that is, a pressure of about 500 milliTorr. These disadvantages include difficulty of doping the resultant film with phosphorus, and controlled-source delivery due to the low vapor pressure of TEOS. Lagendijk also points out that attempts at an all-liquid process to produce borophosphosilicate glass have met with limited success. He further equates the dopant effect within a broad range of phosphorus, boron, antimony, arsenic and chromium compounds, but only when used with silicon compounds having no carbon-oxygen-silicon bonds, and two or more silicon atoms.

In bottle applications, the coatings are applied at such low thicknesses, i.e., about 100 Å, that no iridescence is possible. However, the films are not continuous, and this discontinuity makes them unsuitable for other applications. One solution to the discontinuity is to deposit thicker films of a material with a refractive index closer to that of the article. A mixed metal oxide/silicon oxide material deposited at a significantly more rapid rate than has heretofore been achieved would be desirable, as discussed further hereinbelow.

All the silanes disclosed in the prior art for making mixed metal oxide/silicon dioxide coatings have certain features which make them unsatisfactory for commercial development. Some are very corrosive, flammable, or oxygen-sensitive, and require special handling. Others are not readily available, or

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are too expensive for commercial use. Of the materials which can be used, the biggest problem which limits their commercial development in mixed metal oxide/silicon oxide and/or oxynitride intermediate layers has been that of inadequate deposition rates. When the substrate is flat glass and the deposition process is CVD at ambient pressure, the deposition rate of the intermediate layers must be high enough to coat a production-line glass ribbon traveling at line speeds as high as about 15 meters per minute (m/min). Rates for deposition of the desired layers of about 350 Å are desirable, and rates on the order of 400 to 600 Å/sec are preferable. Such rates have not heretofore been achieved under conditions which permit continuous, mass production of glass with properties.

To overcome the problems as discussed hereinabove, silica precursors are needed which are inexpensive, readily available, easy to handle, and have adequate deposition rates when vaporized with metal oxide precursors. Alkoxysilanes such as TEOS, a commodity chemical, would be desirable. However, prior to the present invention, it has not been possible to deposit silicon oxide films from TEOS by atmospheric-pressure CVD at commercially acceptable deposition rates, except at temperatures at or above 700 degrees Celsius (°C). Some success has been achieved at temperatures of from about 450 to about 680°C, but only by modifying the atmospheric-pressure CVD process through plasma enhancement or reduced pressure, neither of which is generally acceptable for commercial use on a continuous glass ribbon. Additives such as oxygen, ozone, or trimethyl phosphite have also been used in these modified processes, but the rates achieved are still lower than those needed for an effective commercial system.

D. S. Williams and E. A. Dein, in *J. Electrochem. Soc.* 134(3) 657-64 (1987), showed that phosphosilicate and borophosphosilicate glass films with controllable refractive index can be deposited at rates of about 200 Å/sec between 515 and 680°C by the low-pressure CVD of TEOS with phosphorous or boron oxides in concentrations which varied as a function of the additive used. The low-pressure process described here is not amenable to a continuous on-line application of oxides.

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In *Proceedings, 2nd International ULSI Science and Technical Symposium, ECS Proceedings Vol. 98(9), 571-78 (1989)*, D. A. Webb et al. reported that silicon oxide films could be deposited from TEOS at rates of about 125 Å/sec in a plasma-enhanced CVD process using oxygen. However, plasma-enhanced CVD is not a viable option for the continuous commercial application of oxide films to glass, being a batch process requiring complex and costly low-pressure apparatus.

A. K. Hochberg and D. L. O'Meara in *J. Electrochem. Soc. 136(6) 1843 (1989)* reported enhanced deposition of silicon oxide films at 570°C by CVD at low pressure when trimethylphosphite was added to TEOS. As with plasma-enhanced CVD, however, low-pressure CVD is not readily utilized for the continuous commercial application of silicon-oxide films on a moving glass sheet to produce a coated-glass article, due at least in part to the cost and complexity of the device used for deposition at low pressure.

From a review of the prior art, it cannot be determined what precursor combinations, if any, can be used for continuous deposition, under conditions and at a rate suitable for mass production, of mixed metal oxide/silicon oxide films at adequate rates from readily available and relatively inexpensive reagents.

Primary or secondary coatings on glass substrates are further useful to enhance or complement properties of either the substrate or one or more of the coatings thereon, improvement of iridescence being only one application. Other uses of coatings include, e.g., protection of the substrate surface from abrasion, addition of color to clear glass, and screening of particular wavelengths of incident radiation.

DISCUSSION OF THE INVENTION

The present invention is the method of depositing a film onto a glass substrate which comprises depositing at least a first layer in the presence of at least one accelerant. More particularly, the present invention is the CVD method of depositing at atmospheric pressure onto a glass substrate at least one layer of an amorphous coating. This invention provides an article having thereon an improved coating with specific properties such as, e.g., controlled

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refractive index, abrasion resistance, color enhancement, low emissivity, selective light filtration, and anti-iridescence on flat-glass substrates. The method of the invention is CVD at atmospheric pressure and at temperatures lower than 700°C, at rates greater than about 350Å/sec. by using a mixture
5 which includes at least one precursor for a metal oxide, selected from the group consisting of volatile compounds of tin, germanium, titanium, aluminum, zirconium, zinc, indium, cadmium, hafnium, tungsten, vanadium, chromium, molybdenum, iridium, nickel and tantalum. The mixture further includes a precursor for silicon dioxide, and one or more additives selected from the
10 group consisting of phosphites, borates, water, alkyl phosphine, arsine and borane derivatives; PH_3 , AsH_3 and B_2H_6 ; and O_2 , N_2O , NF_3 , NO_2 and CO_2 . The additives are termed "accelerants" herein; the accelerants serve to increase the rate of deposition of the film onto the glass from the mixture. The mixture of precursors and additives is gaseous under the conditions of
15 application required to produce the article; the reaction of the materials in the gaseous mixture with atmospheric or added oxygen provides the corresponding oxides which are deposited on the glass substrate.

Those skilled in the art will understand that precursors and materials discussed in this specification must be sufficiently volatile, alone or with other
20 materials, and sufficiently stable under the conditions of deposition, to be a part of the composition from which the desired films are deposited.

Precursors for deposition of metal oxides include, e.g., aluminum alkyls and alkoxides, cadmium alkyls, germanium halides and alkoxides, indium alkyls, titanium halides, zinc alkyls, and zirconium alkoxides. Specific
25 examples of such compounds include, e.g., $\text{Al}(\text{C}_2\text{H}_5)_3$, CrO_2Cl_2 , GeBr_4 , $\text{Ti}(\text{OC}_3\text{H}_7)_4$, TiCl_4 , TiBr_4 , $\text{Ti}(\text{C}_5\text{H}_7\text{O}_2)_4$, $\text{Zr}(\text{OC}_5\text{H}_{11})_4$, $\text{Ni}(\text{CO})_4$, VCl_4 , $\text{Zn}(\text{CH}_3)_2$, and the like.

Tin precursors include those described by the general formula $\text{R}_n\text{SnX}_{4-n}$, where R is independently chosen from straight, cyclic, or branched-chain alkyl
30 or alkenyl of from one to about six carbons; phenyl, substituted phenyl, or $\text{R}'\text{CH}_2\text{CH}_2-$, where R' is $\text{MeO}_2\text{C}-$, $\text{EtO}_2\text{C}-$, $\text{CH}_3\text{CO}-$, or $\text{HO}_2\text{C}-$; X is selected from the group consisting of halogen, acetate, perfluoroacetate, and their

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mixtures; and where n is 0, 1, or 2. Preferred precursors for tin oxide in the article of this invention are the organotin halides.

Precursors for silicon oxide include those described by the general formula $R_mO_nSi_p$, where m is from 3 to 8, n is from 1 to 4, p is from 1 to 4, and R is independently chosen from hydrogen and acyl, straight, cyclic, or branched-chain alkyl and substituted alkyl or alkenyl of from one to about six carbons, and phenyl or substituted phenyl. Preferred precursors for silicon oxide include tetraethylorthosilicate, diacetoxydi-t-butoxysilane, ethyltriacetoxysilane, methyltriacetoxysilane, methyldiacetoxysilane, tetramethyldisiloxane, tetramethylcyclotetrasiloxane, dipinacoloxysilane, 1,1-dimethylsila-2-oxacyclohexane, tetrakis (1-methoxy-2-propoxy) silane, and triethoxysilane.

Suitable accelerants include phosphite and borate derivatives of the general formula $(R''O)_3P$ and $(R''O)_3B$, where R'' is independently chosen from straight, cyclic, or branched-chain alkyl or alkenyl of from one to about six carbons; phenyl, substituted phenyl, or $R'''CH_2CH_2-$, where R''' is MeO_2C- , EtO_2C- , CH_3CO- , or HO_2C- ; R'' is preferably alkyl or alkenyl of from 1 to 4 carbons in length. Particularly preferred accelerants are those selected from the group consisting of boron and phosphorus esters; most preferred are TEB and TEP.

The precursors to the overcoated layer comprise MBTC or any of the organotins described by the general formula R_nSnX_{4-n} above, and a material chosen to impart a semi-conductive property to the tin oxide; such materials include, e.g., antimony compounds such as trimethylantimony, phosphorous compounds such as triethylphosphine, and fluorine-containing compounds such as trifluoroacetic acid, trifluoroacetic anhydride, ethyl trifluoroacetate, 2,2,2-trifluoroethanol, ethyl 4,4,4-trifluoroacetoacetone, heptafluorobutyryl chloride, and hydrogen fluoride. The tin oxide layer can also be made conductive by depositing sub-stoichiometric films having the composition SnO_{2-x} , wherein x is a non-integer having a value between zero and 1, and wherein the value of x can vary within a given film. The materials for imparting semi-conductive properties to the tin oxide can also be added to the precursors for the first

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layer, to enhance the emissivity of the entire coating system, i.e., the emissivity of the combined first and second layers.

Those skilled in the art will realize that the tin oxide can be replaced in these films entirely or in part by the oxides of other metals such as, e.g., germanium, titanium, aluminum, zirconium, zinc, indium, cadmium, hafnium, tungsten, vanadium, chromium, molybdenum, iridium, nickel and tantalum.

DESCRIPTION OF THE PREFERRED EMBODIMENTS

One embodiment of this invention is the method of producing an article having a film deposited at atmospheric pressure wherein the film comprises one or more mixed metal oxide/silicon dioxide films on a glass substrate, the deposition being made from a mixture comprising a metal oxide precursor, a silicon dioxide precursor, and at least one additive which improves or accelerates the deposition rate significantly when compared to the deposition rate without the additive. The deposited films can contain additional oxides related to the additives used. Further, the deposited mixed oxide films can have specific properties in their own right such as, e.g., designed refractive index, or can be combined with other films, under- or overcoated, or both, to have a combined property such as, e.g., color neutrality or lubricity.

The preferred embodiment of the present invention is the method which comprises the chemical-vapor deposition of a film on a glass substrate at atmospheric pressure at a rate of deposition greater than about $350\text{\AA}/\text{sec.}$, from a precursor of tin oxide and a precursor of silicon oxide, an accelerant selected from the group consisting of organic phosphites, organic borates, water and mixtures thereof, and a source of oxygen, wherein the film comprises at least a first layer of tin oxide and silicon oxide. The gaseous composition is at a temperature below about 200°C at atmospheric pressure.

In a more-preferred embodiment, the mixed metal oxide/silicon dioxide film produced by the method of this invention comprises multiple tin oxide/silicon dioxide layers of, e.g., increasing refractive index; further, a chosen property of a given layer, such as, e.g., the refractive index, can vary continuously such that an overcoated layer of tin oxide will have minimal

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reflected color. A given layer may thus have a concentration of silicon oxide and tin oxide different from the concentrations of silicon oxide and tin oxide in an adjacent layer. The films can also contain oxides of the accelerants, particularly when the additives contain phosphorus or boron.

5 In a most-preferred embodiment of this invention, the precursors to the mixed oxide layer comprise organotin halides generally and monobutyltin trichloride (MBTC) in particular, TEOS, and the accelerant triethyl phosphite (TEP).

10 The compositions of the films produced by this invention were determined by X-ray diffraction (XRD) and X-ray photoelectron spectroscopy (XPS). The article of the present invention is prepared by a process using accelerants whereby the process provides a commercially acceptable continuous CVD deposition of oxide films on moving glass, especially on a modern float-glass line, where the batch processes of the prior art are entirely inapplicable.

15 The effects of added water and added phosphites and borates on the refractive index and deposition rate of TEOS-based mixed oxide films are shown in the following Tables. These results are contrasted to those in Tables III and IV, which show the effect of the additives oxygen and a Lewis acid.

20 Table I shows the effect of added water. As the water concentration is increased, regardless of the tin/silicon ratio or the gas velocity, the deposition rate increases to commercially significant levels. These rate increases are also accompanied by increases in refractive index. In the tables here, the reported deposition rates are approximate with a range of about seven percent, unless
25 the rate is followed by an expressed \pm uncertainty.

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TABLE I

Effect of Water Concentration on Mixed Oxide
Refractive Index and Deposition Rate

5	<u>MBTC</u> mol%	<u>TEOS</u> mol%	<u>Water</u> mol%	<u>R.I.</u>	<u>Dep. Rate</u> <u>Å/sec</u>
665°C glass temperature, 160°C system temperature, 50 l/min gas flow.					
10	0.71	0.71	0.00	1.54	25
	0.71	0.71	0.15	1.73	340
	0.71	0.71	0.24	1.74	400
665°C glass temperature, 160°C system temperature, 12.5 l/min gas flow.					
15	1.05	0.59	0.00	1.74	290
	1.05	0.59	0.60	1.78	330
	1.05	0.59	1.10	1.80	480

20 While 160°C is preferred, the system temperature can be from about 125 to about 200°C.

Table II shows the effects of added TEP and of mixtures of TEP and lower-alkyl borate esters such as triethyl borate (TEB). The results show that TEP is very effective in accelerating the deposition rates of the mixed-oxide films to a high rate at specific and controlled refractive-index values.

25 Additions of TEB at low levels to the TEP resulted in an additional small increase in rate. As used in this specification, the term "high rate," as applied to the film deposition described herein, is greater than about 350Å/sec, and preferably about 400Å/sec or higher. All the films produced
30 under the conditions of Table II were clear.

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TABLE II

Effect of MBTC/TEOS/TEP Concentrations on Deposition Rate

	<u>%TEOS</u>	<u>%MBTC</u>	<u>%TEP</u>	<u>%TEB</u>	<u>R.I.</u>	<u>Dep. Rate</u> <u>Å/sec</u>
5						
	0.80	0.16	-	-	1.69±.02	38± 3
	0.80	0.11	0.76	-	1.58±.01	542± 8
	0.80	0.16	0.76	-	1.60±.01	416±22
10						
	0.78	1.56	0.75	-	1.67±.01	505± 4
	0.78	1.84	0.75	-	1.69±.01	476±45
	0.28	1.56	0.36	-	1.73±.01	231±46
15	0.27	1.56	0.62	-	1.71±.01	381±15
	0.27	1.56	0.75	-	1.70±.01	482± 6
	0.27	1.56	0.75	-	1.70±.01	482±16
	0.27	1.56	0.74	0.18	1.70±.02	492±13
20	0.79	0.16	0.76	0.19	1.59±.01	473±56

25 The glass temperature was 665°C, its speed, 0.56 m/sec; system temperature 160°C, air. MBTC, TEOS, and TEP or the mixture of TEP and TEB were injected separately into the vaporizer section of the coater. Each data point was the average of three samples. The dew point was from -74 to -78°C.

Table III shows the effect of added oxygen. Increasing the oxygen concentration increases the deposition rate significantly, but not to the levels needed for commercial application.

TABLE III

Effect of Oxygen Concentration On Mixed Oxide
Refractive Index and Deposition Rate

	<u>MBTC</u> <u>mol%</u>	<u>TEOS</u> <u>mol%</u>	<u>Oxygen</u> <u>vol% of air</u>	<u>R.I.</u>	<u>Dep. Rate</u> <u>Å/sec</u>
30					
35					
	0.71	0.71	20	1.54	25
	0.71	0.71	50	1.63	50
	0.71	0.71	75	1.65	160
	0.71	0.71	100	1.66	240
40					

665°C glass temperature, 160°C system temperature, 50 l/min gas flow.

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Table IV shows the effect of added Lewis acid, which in this case is excess MBTC. As the concentration increases, the rate increases, although not to the levels needed for commercial application.

Table IV

Effect of MBTC Concentration on Mixed Oxide
Refractive Index and Deposition Rate

	<u>MBTC</u> mol%	<u>TEOS</u> mol%	<u>R.I.</u>	<u>Dep. Rate</u> <u>Å/sec</u>
5				
10	0.48	0.47	1.78	160
	0.48+0.23	0.48	1.78	200
	0.48+0.47	0.47	1.85	300

15 665°C glass temperature, 160°C system temperature, 50 l/min gas flow.

The data in the tables show that effective CVD of mixed oxide films can be achieved at commercial rates by the present invention, with concomitant control of refractive index. The following examples illustrate preferred embodiments of this invention.

Example 1

A square plate of soda-lime silica glass, 9 centimeters (cm.) on a side, was heated on a hot block to 665°C. A gas mixture of about 0.16 mol% MBTC, 0.80 mol% TEOS, 0.75 mol% TEP, and the balance hot air at 160°C was directed over the glass at a rate of 12.5 liters per minute (l/min) for about 10 seconds. The center of the glass surface was uniformly coated with a film which had a pale green color in reflected light. Using the Prism Coupler technique, the refractive index was found to be 1.60 and the thickness was about 4260 Å, corresponding to a deposition rate of about 426 Å/sec. Similarly deposited films have been shown to be amorphous by XRD, and to be composed of oxides of tin, silicon and phosphorus by XPS.

Example 2

A gas mixture of about 1.84 mol% MBTC, 0.78 mol% TEOS, 0.75 mol% TEP, and the balance hot air was directed over a glass surface in the same manner as described in Example 1. The resulting film had a pale magenta color in reflected light. The refractive index was found to be 1.68,

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and the thickness was about 4930 Å, corresponding to a deposition rate of about 493 Å/sec. Similarly deposited films have been shown to be amorphous by XRD, and to be composed of oxides of tin, silicon and phosphorus by XPS.

5

Example 3

A gas mixture of about 1.22 mol% MBTC, 0.58 mol% TEOS, 1.09 mol% H₂O and the balance hot air was directed over a glass surface as described in Example 1, but for eight seconds. The resulting film had a green color in reflected light. The refractive index was found to be 1.78, and the film thickness was about 4650 Å, which corresponds to a deposition rate of about 580 Å/sec. From XRD analysis, similarly deposited films have been found to consist of collapsed tetragonal unit cells of tin oxide, indicating some solid-solution formation with silicon dioxide. XPS analysis shows that the films comprise oxides of tin and silicon.

15

Example 4

Each of the films described in Examples 1 through 3 was successively deposited for one second in ascending-index order. The multi-layer film was then overcoated with about 3200 Å of fluorine-doped tin oxide. This film construction provided a transparent article with essentially no reflected color under conditions of daylight illumination.

20

Example 5

A 9-cm. square of soda-lime silica glass was heated on a hot block to 665°C. A gas mixture of about 1.04 mol% MBTC in air at 160°C, and a gas mixture of 1.04 mol% TEOS and 0.20 mol% TEP in air at 160°C were directed through two microprocessor-controlled globe valves over the glass at a total flow rate of 12.5 l/min for 30 sec. The globe valves were simultaneously opened and closed at a programmed rate such that the gas composition impinging on the glass sample was continuously changed from a mixture of high TEOS/TEP and low MBTC to a mixture of low TEOS/TEP and high MBTC. The center of the glass surface was uniformly coated with a film consisting of oxides of tin, silicon and phosphorus as determined by XPS analysis. As the film thickness increased, the amount of tin gradually

30

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increased, while the amount of silicon and phosphorus decreased. The refractive index was calculated from these data, and from data derived from standard films, and found to lie between 1.52 and 1.87. This film construction provided an article with essentially no reflected color when
5 overcoated with fluorine-doped tin oxide.

Example 6

A gas mixture of about 0.16 mol% MBTC, 0.80 mol% TEOS, and the balance hot air was directed over a glass surface as described in Example 1 for about 60 seconds. The resulting film had a magenta color in reflected
10 light, and a refractive index of 1.69. The film thickness was about 2260 Å, corresponding to a deposition rate of about 38 Å/sec.

Example 7

A 0.5-l clear-glass beverage bottle was rotated and heated to about 600°C in an oven over a three-minute period. The heated bottle was
15 transferred into a coating chamber, where it was contacted with a vapor mixture of 0.16 mol% MBTC, 0.80 mol% TEOS, 0.75 mol% TEP, and the balance hot air at about 170°C for 10 sec. The resulting film was magenta-blue in color, and was uniformly distributed on the sidewalls of the container from shoulder to base. The deposition rate was estimated to be about 200
20 Å/sec from the film color, compared to about 50 Å/sec for the bottle coated only with the vapor mixture of MBTC and TEOS.

From a review of the foregoing tables and examples, those skilled in the art will realize that TEB, TEP, and water serve as accelerants in the CVD of oxide films on glass, and that TEP and TEB are synergistic in
25 accelerating the deposition rate of TEOS and MBTC. Accelerants useful in this invention are chosen from the group consisting of borate and phosphite esters, alkyltin halides, and water.

Modifications and improvements to the preferred forms of the invention disclosed and described herein may occur to those skilled in the art who
30 come to understand the principles and precepts hereof. Accordingly, the scope of the patent to be issued hereon should not be limited solely to the

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embodiments of the invention set forth herein, but rather should be limited only by the advance by which the invention has promoted the art.

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WHAT IS CLAIMED IS:

1. The method which comprises the chemical-vapor deposition of a
5 film on a glass substrate at atmospheric pressure at a rate of deposition
greater than about $350\text{\AA}/\text{sec.}$, from a precursor of tin oxide and a precursor
of silicon oxide, an accelerant selected from the group consisting of organic
phosphites, organic borates, water and mixtures thereof, and a source of
oxygen, wherein the film comprises at least a first layer of tin oxide and
10 silicon oxide.
2. The method of claim 1 wherein the substrate is transparent flat
glass at a temperature above about 450°C.
3. The method of claim 1 wherein the substrate is transparent flat
glass at a temperature of from about 450 to about 650°C.
- 15 4. The method of claim 1 wherein the glass with the deposited film
has essentially no reflected color in daylight.
5. The method of claim 1 wherein the glass substrate is moving and
the deposition is continuous.
6. The method of claim 1 wherein the precursors are volatile.
- 20 7. The method of claim 1 wherein the first layer comprises tin
oxide, silicon oxide, and phosphorus oxide.
8. The method of claim 1 wherein the first layer comprises tin,
silicon, phosphorus and boron oxides.
9. The method of claim 1 wherein the accelerant is triethyl
25 phosphite.
10. The method of claim 1 wherein the precursor of the tin oxide is
 $\text{R}_n\text{SnX}_{4-n}$, where R is a straight, cyclic, or branched-chain alkyl, or alkenyl of
from one to about six carbons; phenyl, substituted phenyl, or $\text{R}'\text{CH}_2\text{CH}_2-$,
where R' is $\text{MeO}_2\text{C-}$, $\text{EtO}_2\text{C-}$, $\text{CH}_3\text{CO-}$, or $\text{HO}_2\text{C-}$; X is selected from the
30 group consisting of halogen, acetate, perfluoroacetate, and their mixtures; and
where n is 0, 1, or 2.

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11. The method of claim 1 wherein the precursor of the tin oxide is an alkyltin halide.

12. The method of claim 1 wherein the precursor of the tin oxide is an alkyltin chloride.

5 13. The method of claim 1 wherein the precursor of the tin oxide is chosen from the group consisting of monobutyltin trichloride, dibutyltin dichloride, tributyltin chloride, and tin tetrachloride.

14. The method of claim 1 wherein the precursor of silicon oxide is $R_mO_nSi_p$, where m is from 3 to 8, n is from 1 to 4, p is from 1 to 4, and R is independently chosen from hydrogen and acyl, straight, cyclic, or branched-chain alkyl and substituted alkyl or alkenyl of from one to about six carbons, and phenyl or substituted phenyl.

15 15. The method of claim 1 wherein the precursor of silicon oxide is selected from the group consisting of tetraethylorthosilicate, diacetoxydi-t-butoxysilane, ethyltriacetoxysilane, methyltriacetoxysilane, methyldiacetoxysilane, tetramethyldisiloxane, tetramethylcyclotetrasiloxane, dipinacoloxysilane, 1,1-dimethylsila-2-oxacyclohexane, tetrakis (1-methoxy-2-propoxy) silane, and triethoxysilane.

20 16. The method of claim 1 wherein the precursor of silicon oxide is tetraethylorthosilicate.

17. The method of claim 1 wherein the accelerant comprises triethyl phosphite.

18. The method of claim 1 wherein the accelerant comprises triethyl phosphite and triethyl borate.

25 19. The method of claim 1 wherein the rate of deposition is greater than about 400Å/sec.

20. The method of claim 1 wherein the first layer is amorphous.

21. The method of claim 1 wherein the first layer comprises a plurality of layers, and at least a second layer is deposited on the first layer.

30 22. The method of claim 21 wherein the second layer comprises a tin oxide.

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23. The method of claim 21 wherein the second layer comprises a mixture of tin oxide and a fluorine compound.

24. The method of claim 21 comprising a plurality of layers, each layer having a separate refractive index.

5 25. The method of claim 21 wherein the first layer has a concentration of silicon oxide and tin oxide different from the concentrations of silicon oxide and tin oxide in an adjacent layer.

26. The method of claim 21 wherein the first layer has a refractive index which changes continuously between the substrate and the second layer.

10 27. The method of claim 21 wherein the second layer comprises a doped tin oxide.

28. The method of claim 21 wherein the second layer is deposited from a precursor mixture comprising monobutyltin trichloride and a fluorine-containing material.

15 29. The method of claim 21 wherein the first layer is deposited from a precursor mixture comprising monobutyltin trichloride and tetraethyl orthosilicate in the presence of triethyl phosphite.

20 30. The method for producing on glass a layer having a controlled index of refraction, which comprises depositing at least one amorphous layer on the glass at a rate greater than about 400Å/sec., applying to the glass a mixture of a tin oxide precursor, a silicon oxide precursor, and at least one accelerant chosen from the group consisting of boron and phosphorus esters and water.

25 31. The method of claim 30 comprising continuous chemical-vapor deposition of a mixture of monobutyltin trichloride, tetraethylorthosilicate and an accelerant onto a moving glass sheet, wherein the glass is at a temperature of from about 450 to about 650°C.

INTERNATIONAL SEARCH REPORT

International application N .

PCT/US92/10872

A. CLASSIFICATION OF SUBJECT MATTER

IPC(5) :B05D 5/06; C23C 16/22,16/40

US CL :427/248.1,255.255.1,255.3,255.7,166

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

U.S. : 427/248.1, 255.1, 255.3, 255.7, 166

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practicable, search terms used)

APS Search

phosphite ester#, triethylphosphite#, trimethyl phosphite, oxide #; deposition rate rate; glass substrate

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X Y	US,A, 4,880,664 (O'DOWD ET AL) 14 NOVEMBER 1989 See col. 3, line 67- col. 4 line 8, col. 4 lines 18-45 col. 5, lines 44-48 col. 6 lines 17-20	<u>1.7-8,23-24,27</u> 2-6, 9-22, 25- 26,28-31
Y	US,A, 4,308,316 (GORDON) 29 DECEMBER 1984 See col. 4 line 24 col. 9 lines 4-8 col. 5 lines 27-29	1-31
Y	US,A, 5,028,566 (LAGENDIJK) 02 JULY 1991 See col. 4 lines 41-42	1-31

☒ Further documents are listed in the continuation of Box C. ☐ See patent family annex.

* Special categories of cited documents:	*T	later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention
A document defining the general state of the art which is not considered to be part of particular relevance	*X*	document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone
E earlier document published on or after the international filing date	*Y*	document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art
L document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)	*A*	document member of the same patent family
O document referring to an oral disclosure, use, exhibition or other means		
P document published prior to the international filing date but later than the priority date claimed		

Date of the actual completion of the international search

18 FEBRUARY 1993

Date of mailing of the international search report

16 MAR 1993

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INTERNATIONAL SEARCH REPORT

International application N .
PCT/US92/10872

C (Continuation). DOCUMENTS CONSIDERED TO BE RELEVANT

Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
A,P	US,A, 5,108,479 (HIRANO) 28 APRIL 1992 See col. 7 lines 3-63	1-31
A	US,A, 3,984,581 (DOBLER ET AL) 05 OCTOBER 1976 See col. 2 lines 3-40	1-31